

Sintering behaviour of spinel–alumina composites

SOUMEN PAL, A K BANDYOPADHYAY, P G PAL*, S MUKHERJEE[†] and B N SAMADDAR^{††}

Department of Ceramic Technology, Govt. College of Engineering & Ceramic Technology, Kolkata 700 010, India

[†]School of Material Science & Technology, ^{††}Ex-emeritus Professor (AICTE),

Department of Metallurgical & Material Engineering, Jadavpur University, Kolkata 700 032, India

MS received 21 July 2008

Abstract. Study of alumina–magnesia binary phase diagram reveals that around 40–50 wt% alumina dissolves in spinel (MgAl_2O_4) at 1600°C. Solid solubility of alumina in spinel decreases rapidly with decreasing temperature, which causes exsolution of alumina from spinel phase. Previous work of one of the authors revealed that the exsolution of alumina makes some interlocking structures in between alumina and spinel phases. In the present investigation, refractory grade calcined alumina and spinel powder were used to make different batch compositions. Green pellets, formed at a pressure of 1550 kg cm^{-2} were fired at different temperatures of 1500°, 1550° and 1600°C for 2 h soaking time. Bulk density, percent apparent porosity, firing shrinkage etc were measured at each temperature. Sintering results were analysed to understand the mechanism of spinel–alumina interactions. SEM study of present samples does not reveal the distinct precipitation of needle shaped α -alumina from spinel, but has some effect on densification process of spinel–alumina composites. Microstructural differences between present and previous work suggest an ample scope of further work in spinel–alumina composites.

Keywords. Sintering; stoichiometric spinel; spinel–alumina composites; solid-solution.

1. Introduction

It is known for quite a long time that alumina–magnesia spinel with varying amounts of alumina and magnesia (either alumina-rich or magnesia-rich side) is a very important refractory material for various interesting applications in metallurgical and cement industries. This is due to the fact that the dense alumina–magnesia spinel has adequate ‘hot strength’, high ‘corrosion resistance’ to both acidic and basic slags, and also has good erosion resistance at elevated temperatures. The solid solubility of magnesia in alumina is very low (100 ppm at 1550°C), but its effect on the densification of alumina is remarkable. An addition of small amount (0.2 wt%) of MgO in alumina stops secondary recrystallization of alumina and thereby helps alumina to attain its ‘near theoretical density’. This fact has been interpreted by many workers (Coble 1961; Hener 1979; Burke *et al* 1980; Greskovitch and Anthony Brewer 2001) in various ways. Effect of spinel on sintering kinetics of alumina was studied by two of the authors (Paul and Samaddar 1985; Paul 1986) and showed that excess spinel beyond its solid solubility limit in alumina exists as second phase (spinel). The work also revealed that α -alumina exsolved out as needle shaped precipitate forming inter-

locking structures. The latest Al_2O_3 –MgO binary phase diagram (Kingery *et al* 1976) differs from that reported by Kingery in his first edition (Kingery 1960) in terms of solid solubility of MgO in MgAl_2O_4 spinel.

According to the latest diagram, MgO has considerable solubility in MgAl_2O_4 , and hence magnesia-enriched spinel formed at high temperature is expected to precipitate MgO from spinel. The microstructural evidence of precipitation of MgO in spinel is rarely available in the literature. The incorporation of a second phase in alumina improves the functional properties of alumina by altering the strength (creep rate), microstructure, deformation mechanism and fracture behaviour (Takigawa *et al* 1998; Kim *et al* 2001). Shiono *et al* (2000) also studied the deformation behaviour of fine grained magnesium aluminate spinel prepared by chemical route.

Structurally speaking, most stable form of alumina (α -alumina) has corundum structure and it is stable up to its melting point. The structure is based on hexagonal closed packing of oxygen ions in which 2/3rd of the octahedral sites are filled with aluminium ions. In spinel, the oxygen ions are cubically closed packed creating octahedral and tetrahedral voids. In such sites, 1/8th of the tetrahedral and 1/2 of the octahedral voids are filled by Mg^{2+} and Al^{3+} ions, respectively. In spinel–alumina interaction at high temperatures, Al_2O_3 goes into solid solution in spinel by replacing three Mg^{2+} ions by two Al^{3+} ions. Ulti-

*Author for correspondence (pgp1951@rediffmail.com)

mately, the end product is γ -alumina, which is also cubic in structure.

A combination of magnesia–spinel and alumina–spinel composites with a moderate range of porosity is used as high temperature structural materials. The refractory grade alumina-rich spinels containing up to 90 wt% alumina are commercially used in high alumina castables. The solid solubility of Al_2O_3 in spinel and its subsequent precipitation during actual operation play the vital role for improving the performance of alumina–spinel castables.

Therefore, in the present investigation, the interaction between alumina-enriched spinel and alumina was studied in the entire range of spinel–alumina composites. High temperature interaction between spinel and alumina has been expressed in terms of changes of different physical properties such as bulk density, percent apparent porosity, densification rate, etc with composition of the composites. The SEM results were analysed and also compared with previous work in order to understand the mechanism of interaction.

2. Experimental

2.1 Sample preparation

The refractory grade calcined alumina (with particle size below 15 μm and purity, 99.6%) and alumina-enriched spinel [–300 B.S. sieve and purity ($\text{Al}_2\text{O}_3 + \text{MgO}$) = 99%] were used for the present study. Both the raw materials were characterized in terms of chemistry by conventional wet chemical method (table 1) and particle size distribution by particle size analyser (table 2). Different batch compositions containing spinel and calcined alumina were prepared by mixing in an agate mortar. In each case, the batch size was 60 g and the mixing time was 30 min. The details of batch compositions are given in table 3.

A cylindrical high carbon steel mould of internal diameter 2.55 cm (approx.) was used for moulding green pellets at a forming pressure of 1550 kg cm^{-2} with the addition of appropriate amounts of organic binder and water. For each batch composition, six pellets were prepared. The ‘green density’ of pellets in all the batch compositions was measured from dimensions (diameter and thickness) and mass of the dried pellets (i.e. at 110°C for 24 h).

Table 1. Chemical analysis of refractory grade calcined alumina and spinel (MgAl_2O_4) powder [loss free basis].

Constituents	Calcined Al_2O_3 (wt%)	MgAl_2O_4 (wt%)
Al_2O_3	99.50	77.40
MgO	Traces	21.60
Na_2O	0.35	0.28
Fe_2O_3	0.06	0.12
SiO_2	0.04	0.04
CaO	Traces	0.28

2.2 Sintering

Three temperatures, viz. 1500°, 1550° and 1600°C, were chosen for sintering studies. Two green pellets of each batch were sintered at each temperature for an averaging procedure in order to have a statistical reliability of the data. The same heating schedule was followed as: 10°C min^{-1} from room temperature to 1000°C, and 5°C min^{-1} from 1000°C to the respective sintering temperature. A soaking time of 2 h was given for sintering at each temperature. In each case, the furnace was forced to cool down to 800°C, and then normal cooling was followed down to room temperature.

2.3 Physical properties

The density of the pellets sintered at different temperatures was measured from dimensions and mass of the pellets. The apparent porosity and bulk density of the pellets were measured by boiling method. There was hardly any difference between the measured values of densities by two methods. The firing shrinkage and percent densification during sintering were calculated from dimensions and density of green and fired pellets. Pycnometric densities of all the composites (A1 to A9) sintered at 1600°C were measured by following standard method.

Table 2. Particle size distribution of refractory grade calcined alumina and spinel.

Particle size (μm)	Materials	
	Calcined Al_2O_3 (wt% under)	MgAl_2O_4 (wt% under)
53.5	100.0	97.5
37.6	100.0	77.8
28.1	100.0	41.2
21.5	100.0	20.5
16.7	100.0	13.2
13.0	98.4	5.6
10.1	86.5	1.2
7.9	69.7	
4.8	33.9	
3.0	7.7	
1.9	1.9	

Table 3. Batch compositions (batch size 60 g).

Batch number	Calcined Al_2O_3 (wt%)	MgAl_2O_4 (wt%)
A1	0	100
A2	5	95
A3	10	90
A4	25	75
A5	50	50
A6	75	25
A7	90	10
A8	96	4
A9	100	0

Table 4. Chemical compositions of different batches of spinel–alumina composites.

Batch number	Chemistry		
	Al ₂ O ₃ (wt%)	MgO (wt%)	Impurities (wt%) (Fe ₂ O ₃ + Na ₂ O + SiO ₂ + CaO)
A1	77.40	21.60	1.00
A2	78.51	20.52	0.97
A3	79.61	19.44	0.95
A4	82.92	16.20	0.88
A5	88.45	10.80	0.75
A6	93.98	5.40	0.62
A7	97.29	2.16	0.55
A8	98.62	0.86	0.52
A9	99.50	Nil	0.50

Table 5. Green density (dried at 110°C, 24 h) of pellets of different batch compositions (forming pressure = 1550 kg cm⁻²).

Batch number	Green densities (g cm ⁻²)		
	Maximum	Minimum	Average
A1	2.48	2.45	2.45
A2	2.48	2.45	2.46
A3	2.45	2.45	2.45
A4	2.46	2.43	2.44
A5	2.44	2.39	2.41
A6	2.41	2.37	2.39
A7	2.40	2.34	2.35
A8	2.30	2.26	2.29
A9	2.26	2.23	2.24

2.4 Microstructural studies

Scanning electron microscopic (SEM) studies were conducted on fractured surfaces of the samples (A1 to A8, sintered at 1600°C for 2 h) coated with gold in JEOL-5200 scanning electron microscope operated at 15 kV.

3. Results and discussion

3.1 Characterization of powders

The chemical analyses of calcined alumina and spinel powder (table 1) indicates 99.5 and 99% purity, respectively. It is expected that sintering of either spinel or alumina or combinations of spinel–alumina admixture will be accompanied by solid state sintering. Their particle size distribution (table 2) reveals that fineness of powders are in sub-sieve but not sub-micron range, and they will not be so much interactive to give enhanced densification (near theoretical) at 1500–1600°C temperature range.

The chemical compositions of all the batches containing alumina and spinel were calculated (table 4) based on chemistry of calcined alumina and spinel (table 1). The chemistry of spinel–alumina composites (A1 to A8) reveals that each composite is a combination of spinel and

corundum (α -alumina) phase. The green densities of spinel and alumina powders formed at identical forming pressure (1550 kg cm⁻²) are 2.45 and 2.24 g cm⁻³, respectively although alumina has a higher specific gravity (3.95) than spinel (3.638). It is pertinent to note that the lower green density may be due to the presence of considerable amount of ‘inter-granular open voids’ (not deformable during forming) in calcined alumina as compared to spinel. This results in gradual reduction of green density with increase in alumina content in the ‘spinel compositions’ (table 5). True specific gravity of each batch composition of sintered pellets (1600°C, 2 h) was measured by pycnometric method and reported in table 6.

3.2 Physical characteristics

For better understanding of the mechanism of spinel–alumina interaction at different temperatures (1500°, 1550° and 1600°C), different parameters related to sintering, viz. fired density, percent apparent porosity, percent firing shrinkage (i.e. volume shrinkage), percent closed pores and percent densification with respect to green density were plotted against wt% alumina that was added in the ‘spinel composite’.

The spinel powder used for the present investigation contained around 22 wt% α -alumina and 77 wt% stoichiometric magnesium aluminate spinel. Phase compositions of composites were calculated in terms of stoichiometric magnesium aluminate spinel and excess α -alumina from chemical composition of composites. Phase compositions and their correlation with % open pores, closed pores and apparent specific gravity of composites have been specified in table 7. An addition of 5–10 wt% alumina in spinel causes considerable increase in sintered bulk density at all sintering temperatures (figure 1).

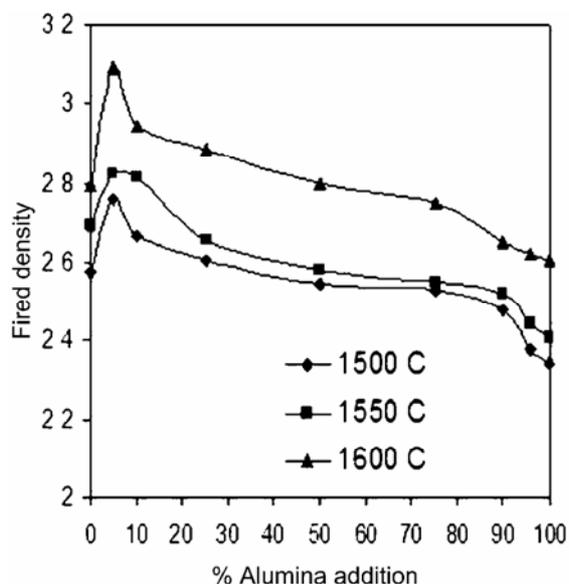
Further addition of alumina causes gradual decrease in the bulk density, which is unexpected since true specific gravity of alumina is greater than spinel. The sintered density of pellets made of only calcined alumina gives lower value, which reveals that calcined alumina is not so

Table 6. True specific gravity and % true porosity of different spinel–alumina composites.

Batch number	True specific gravity (TSG)	% True porosity (100(TSG–BD)/TSG)		
		1500°C, 2 h	1550°C, 2 h	1600°C, 2 h
A1	3.6380	29.08	26.06	23.03
A2	3.6610	24.61	22.70	15.32
A3	3.6762	27.37	23.29	19.75
A4	3.7218	30.14	28.53	22.62
A5	3.7979	32.86	32.07	26.27
A6	3.8739	34.69	34.17	29.01
A7	3.9196	36.73	35.71	32.39
A8	3.9378	39.56	38.04	33.47
A9	3.9500	40.76	38.99	33.92

Table 7. Relation among phase composition and % open pores, % closed pores and apparent specific gravity of spinel–alumina composites at 1600°C for 2 h soaking time.

Batch no.	Phase composition (wt%)		Apparent specific gravity	% Open pores	% Closed pores
	Stoichiometric spinel	Excess α -alumina			
A1	76.68	22.32	3.542	20.94	2.26
A2	72.91	26.19	3.55	12.68	2.64
A3	69.01	30.04	3.445	14.37	5.38
A4	57.51	41.31	3.516	18.1	4.52
A5	38.34	60.91	3.565	21.45	4.83
A6	19.17	80.21	3.606	23.73	5.28
A7	7.67	91.78	3.727	28.89	3.5
A8	3.05	96.43	3.784	30.76	2.71
A9	–	99.5	3.835	31.94	2.91

**Figure 1.** Variation of fired density of spinel–alumina composites with addition of alumina at different temperatures (2 h soaking time).

reactive by itself. Since true specific gravity of different batches gradually increases with addition of alumina, it is expected that the marginal decrease in bulk density will be reflected to a greater extent in an increase of percent apparent porosity (%AP) and percent true porosity (%TP).

The addition of alumina to a certain extent (5–10 wt%) in spinel causes decrease in both %AP (figure 2) and %TP (table 6) of spinel–alumina composites. The addition of 5–10 wt% alumina causes maximum reduction both in %AP and %TP of composites at all temperatures from 1500–1600°C. At 1600°C, the porosities (both %AP and %TP) of the composites are lower than spinel-based composite up to 25 wt% addition of alumina. At higher wt%, addition of alumina, the porosity gradually increases with the amount of addition of alumina. Figure 3 shows the variation of percent closed pores with wt% alumina addition. It is observed that all the composites have closed pores at all the temperatures, and its value increases in the composition range having 10–90 wt% addition of alumina.

Composites containing around 70 wt% to around 20 wt% spinel have higher percentage of closed pores (table 7). Solid solution of alumina in spinel at 1600°C and its precipitation during cooling are complex phenomena. This may be due to complete precipitation of α -alumina from spinel structure in the above composition range, which will cause increase of both open and closed pores. The variations of percent volume shrinkage and percent densification (with respect to green density) with wt% alumina (figures 4 and 5) show the same trend, as is observed in the case of bulk density vs wt% alumina plot.

3.3 Microstructural studies

Microstructural inhomogeneity was observed in SEM images (figure 6) of most of the composites made by the interaction of spinel and alumina at 1600°C. Evidence of void formation due to acicular inhomogeneity in synthetic spinel was studied by Arnold (1960). Here composite A1 is made of admixture of stoichiometric spinel and

alumina. Dissolution of alumina from spinel lattice structure is dependent on thermal cycle of spinel–alumina interaction. Dual microstructure is seen even in spinel-based composite (figure 6A1) where some spinel grain surfaces are clean and most are roughened by precipitation of α -alumina from spinel lattice. The extent of surface roughening due to precipitation of fine alumina gradually increases up to 25 wt% addition of alumina in

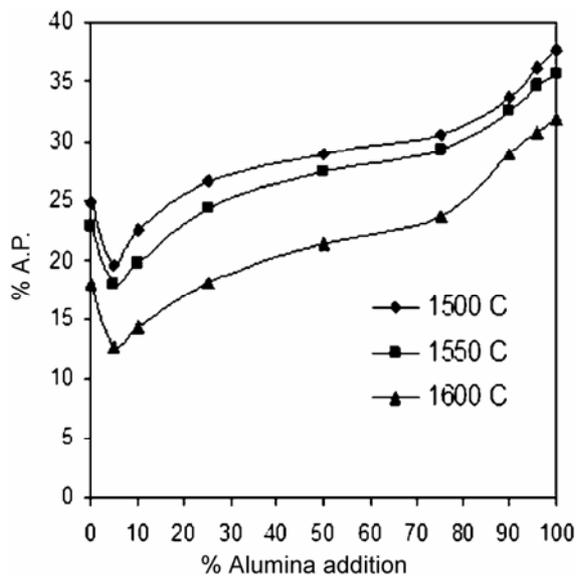


Figure 2. Variation of apparent porosity (AP) of spinel–alumina composites with addition of alumina at different temperatures (2 h soaking time).

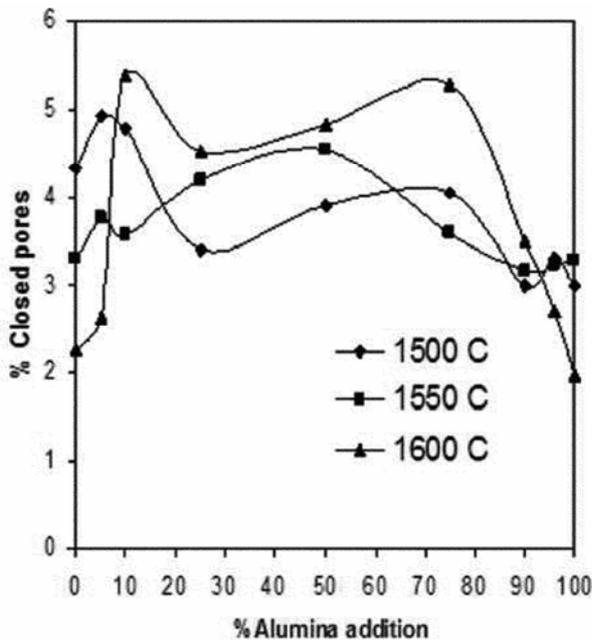


Figure 3. Variation of % closed pores of spinel–alumina composites with addition of alumina at different temperatures (2 h soaking time).

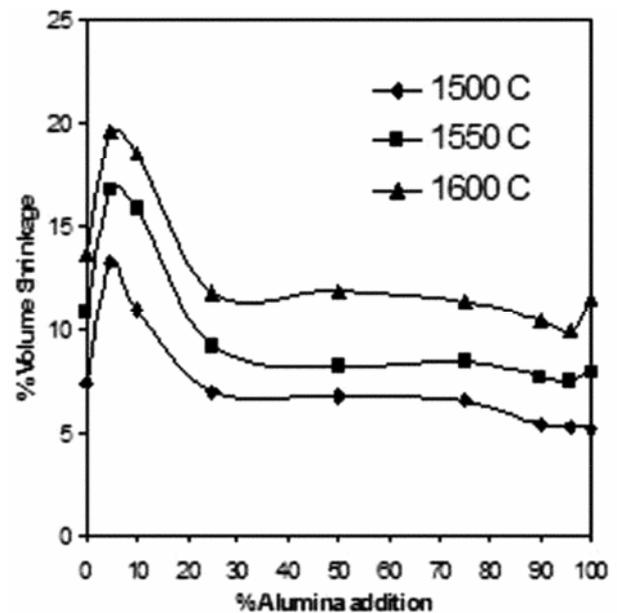


Figure 4. Variation of volume shrinkage of spinel–alumina composites with addition of alumina at different temperatures (2 h soaking time).

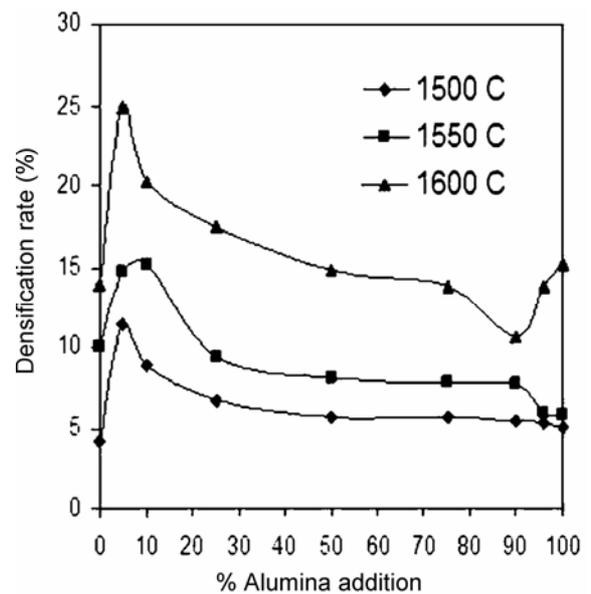


Figure 5. Variation of densification rate (%) of spinel–alumina composites with addition of alumina at different temperatures (2 h soaking time).

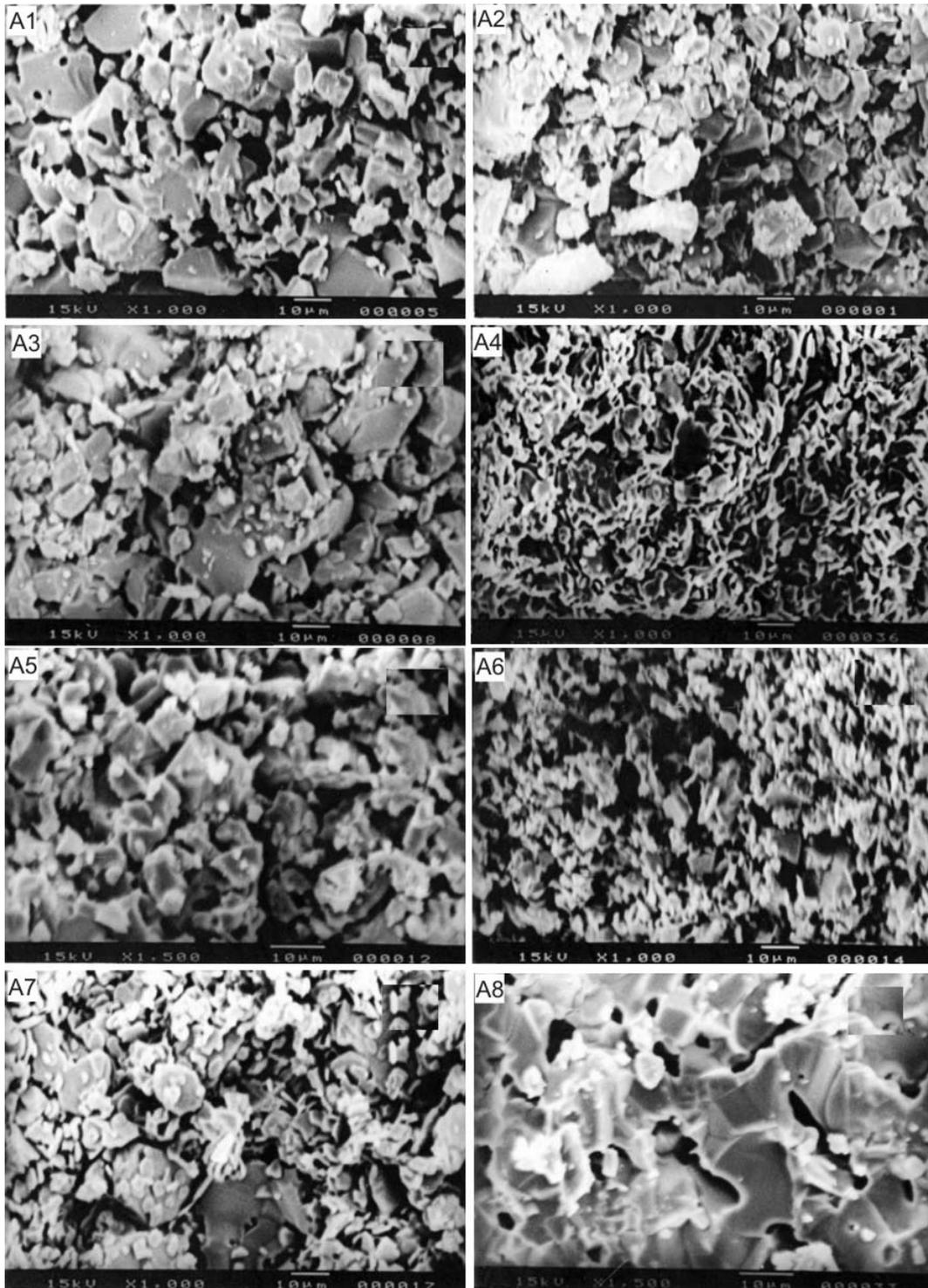


Figure 6. SEM photographs of fractured surface of spinel–alumina composites (figures 6A1–6A8) sintered at 1600°C for 2 h. (Here, A1 to A8 refer to the batch compositions as specified in tables 3–7).

spinel (figures 6A2–6A4). On the other hand, addition of spinel in alumina has positive effects in improvement of physical properties such as increase in bulk density and decrease in porosity, etc of alumina–spinel composites. In the present work, alumina–spinel composites do not have

distinct precipitation of α -alumina and lack development of interlocking structures. Previous work of one of the authors (Paul 1986) may be reproduced (figures 7a–d) to show the development of interlocking structures when submicron mono-size (0.3–0.4 μm) alumina powder

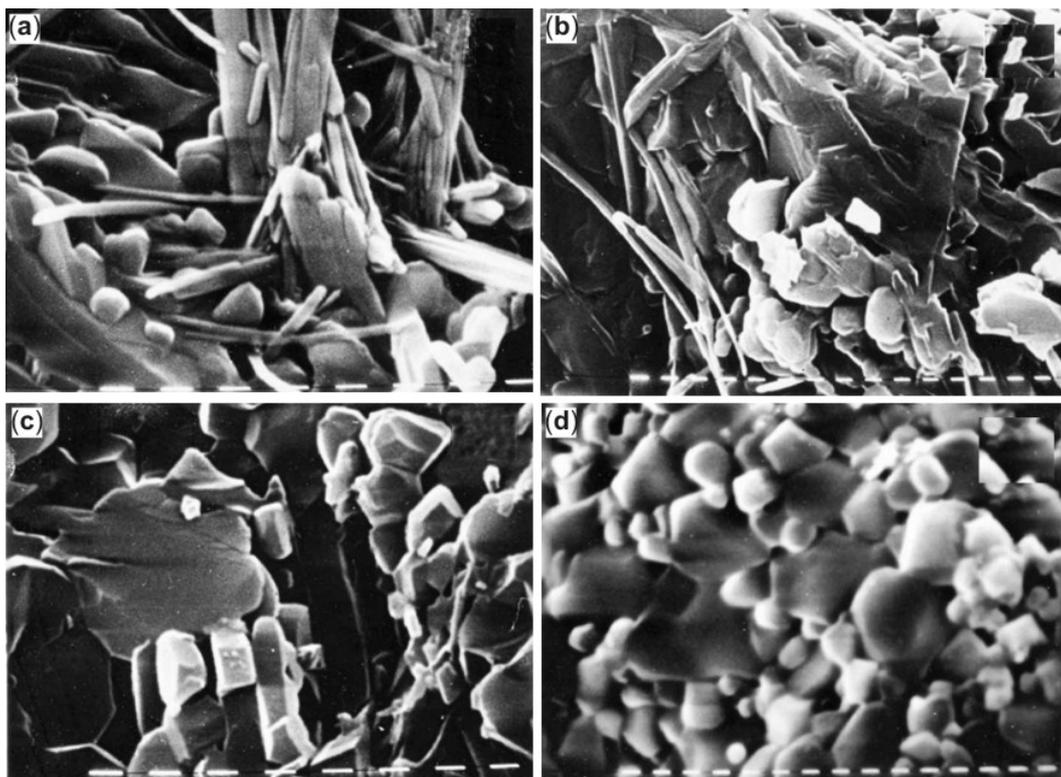


Figure 7. SEM photographs of fractured surfaces of alumina–spinel composites sintered at 1600°C, 2 h soaking (each dash indicates 1 μm): (a) 85 wt% alumina and 15 wt% spinel, (b) 75 wt% alumina and 25 wt% spinel, (c) 50 wt% alumina and 50 wt% spinel, and (d) 100 wt% spinel (reprinted with permission from Paul 1986).

interacted with chemically prepared phase pure magnesium aluminate spinel. Distinct needle form precipitation of α -alumina (figures 7a and b) can be seen when addition of spinel is in the 15–25 wt% range. However, in case of 50–50 wt% spinel–alumina composite, this needle like precipitation of α -alumina is not visible and both spinel and alumina grains are distinctly noticeable in the microstructure. Phase pure sintered spinel (1600°C, 2 h) shows cubical morphology of spinel (figure 7d). In the present work, abnormal grain growth in sintered alumina and distinct precipitation of α -alumina in the composites are not noticeably visible mostly due to coarser granulometry of both alumina and spinel powder as compared to the finer granulometry of the previous work.

Further dissolution of alumina in spinel during heat treatment and subsequent precipitation of α -alumina during cooling are complex processes. This results in generation of both open and closed pores. Also, creation of open pores may be due to stress exerted by precipitated alumina at grain boundary/matrix. Voids created by precipitation process of α -alumina from spinel structure are not completely opened up. It happened mostly in the composites A_3 to A_6 causing higher percentage of closed pores. True specific gravity of spinel–alumina composite depends on

its phase composition as evidenced by true specific gravity of the composites (table 6). Suwa *et al* (1986) studied the interaction between chemically pure alumina and magnesia (93 : 7 wt ratio) gel and observed that the complete phase separation of α -alumina and spinel took place from composite containing γ -alumina and spinel admixture in the form of solid solution at 1100°C for 100 min soaking time.

Literature (Rinne 1928) is available where it has been shown that spinel having higher stoichiometry of alumina (92.7 wt% alumina) has marginally higher specific gravity (3.624) than stoichiometric spinel (3.578), and this behaviour is absolutely abnormal. It is only possible if dissolved alumina is retained in spinel lattice structure at room temperature. SEM photographs (figures 6A4–6A6) of present study having the compositions of spinel: alumina = 60 : 40 to 20 : 80 (table 7) have needle like precipitation of α -alumina. Morphology of composite having spinel: alumina (approximate wt ratio 40 : 60) (figure 6A5) is different from that of the spinel: alumina having a wt ratio of 60 : 40 (figure 6A4) or 20 : 80 (figure 6A6). Secondary grain growth is observed in alumina body (figure 6A8), containing around 3 wt% spinel (table 7), which is also abnormal. This may be due to heterogeneous incorporation of spinel in alumina.

4. Conclusions

The solid solution and precipitation mechanism due to spinel–alumina interaction may be used as a basic model to develop spinel–alumina or alumina–spinel composites. The spinel can retain some amount of alumina in its host structure even at room temperature, which may be evolved as α -alumina during use, thus forming an interlocking structure between two types of (alumina and spinel) grains. This type of microstructural feature is consistent with the trends in the improvement of physical properties such as %AP, BD, %TP, etc of spinel by the addition of alumina, as observed in the present investigation. Further work is in progress to develop spinel based dense composites with the addition of alumina for various industrial applications.

References

- Arnold H 1960 *Z. Krist.* **114** 23
 Burke J E, Lay K W and Prochazka S 1980 *Sintering process sci. res.* (ed.) Kuczynski (New York and London: G C Plenum Press) **13** p. 417
 Coble R L 1961 *J. Appl. Phys.* **32** 793
 Greskovitch C and Anthony Brewer J 2001 *J. Am. Ceram. Soc.* **84** 420
 Hener A H 1979 *J. Am. Ceram. Soc.* **62** 317
 Kim B N, Hiraga K, Morita K and Sakka Y 2001 *Acta Mater.* **49** 887
 Kingery W D, Bowen H K and Uhlmann D R 1976 *Introduction to ceramics* (New York: John Wiley & Sons Inc.) 2nd ed. pp. 131–133
 Kingery W D 1960 *Introduction to ceramics* (New York: John Wiley & Sons Inc.) p. 264
 Paul P G 1986 *Study of sintering kinetics of alumina in presence of magnesio–aluminate spinel*, Ph D Thesis, University of Calcutta, Kolkata
 Paul P G and Samaddar B N 1985 *Trans. Ind. Ceram. Soc.* **44** 132
 Rinne F 1928 *Neues Jahrb. Mineral Geol Beib Bd.* **A58** 43
 Shiono T, Ishitomi H, Okamoto Y and Nishida T 2000 *J. Am. Ceram. Soc.* **83** 645
 Suwa Y, Roy R and Komarneni S 1986 *Mater. Sci. & Eng.* **83** 151
 Takigawa Y, Yoshizawa Y and Sakume T 1998 *Ceram. Int.* **24** 61